

Covalent Bond-

Diatomic-

List the "lucky seven " diatomic molecules-

Molecular formula-

Ionic compound	Covalent compound	
		Compare melting points
		Compare boiling points

Bonding pair of electrons-

Lone pair(non bonding electrons)-

The number of bonds is-

Coordinate covalent bond-

Show the resonance structure of ozone-

Draw Lewis dot diagrams for the following elements: I, S, P, C, B, Li, Mg

Octet rule-exceptions are→

Sigma bond-

Pi bond-

Polar means-

Intermolecular forces-

Two types of Van der waals forces are 1- _____ 2- _____
 Hydrogen Bonds are-

Three types of intermolecular forces arranged from weakest to strongest are
 Weakest-----→strongest
 Dispersion forces→ _____ → _____

Stoichiometry-

Write the equation below using 1- moles 2- atoms 3-molecules 4- volume
 and 5-mass.

Equation: $2\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{H}_2\text{O}(\text{g})$

1)moles

2)atoms

3)molecules

4)volume

5) mass

The steps for doing a mass-mass stoichiometry question are:	
1	
2	
3	
4	

1. Calculate the mass of carbon dioxide produced by the complete combustion of 692 g of C_8H_{18} .

2. In the reaction of silver chloride with zinc nitrate, 15.0 g of silver chloride reacts. How many grams of zinc chloride will form?

Limiting reagent:

Excess reagent:

Formula for percent yield:

VSEPR(What is it?)-

G. N. Lewis-

Make and fill in a table like you see below

Formula and name of Shape	#bonding pairs	#lone pairs (or non-bonding pairs)	Example (and sketch)	Polarity
AX Linear				
	2	0	CO ₂ O=C=O HCN H-C * N	nonpolar polar
AX ₆				
AX ₃ pyramidal				
AX ₂ v-shaped				

*Triple bond

Do p 248 #65a,c,d,68,69 p 249 #82, p 382 #81, p 375 #32(ans 57.7%) p 380 #52 a (hint: do a mass/mass stoichiometry calculation first, to find mass of SiC produced, then do % calculation ans 96.4 %)